Name STURMAN key

Period ____ Date ____

Chapter 8 Crossword

Across

1.) The 2 weakest attractions between molecules.

3.) Several atomic orbitals mix to form the same total number of equivalent hybrid orbitals.

4.) Covalent bond between atoms in which the electrons are shared unequally.

5.) Energy required to break the bond between 2 covalently bonded atoms.

7.) Bond formed by sharing 3 pairs of electrons.

8.)Atoms held together by sharing electrons are joined by a ___.

9.) Represents the covalent bonds by dashes and shows the arrangement of covalently bonded atoms.

10.) Valence-shell electron-pair repulsion theory.

13.) Neutral group of atoms joined together by covalent bonds.

Lame→ 14.) Molecular orbital that can be occupied by 2 electrons of a covalent bond.

15.) Molecule consisting of 2 atoms.

18.) Covalent bond in which the electrons are shared equally by 2 atoms.

19.) Molecule in which one side of the molecule is slightly negative and the opposite side is slightly positive.

21.) The weakest of all molecular interactions.

OMIT → 25.) A compound composed of molecules.

Down

2.) 2 atoms held together by sharing a pair of electrons are joined by____?

3.) Attractive forces in which hydrogen covalently bonded to a very electronegative atom is also weakly bonded to an unshared electron pair of another electronegative atom.

4.)Tightly bound group of atoms that have a positive or negative charge and behaves as a unit.

6.) Chemical formula of a molecular compound.

9.) When 2 atomic orbitals combine to form a molecular orbital that is symmetrical around the axis connecting 2 atomic nuclei, a ____ is formed.

11.) A covalent bond in which the bonding electrons are most likely to be found in sausage-shaped regions above and below the bond axis of the bond atoms.

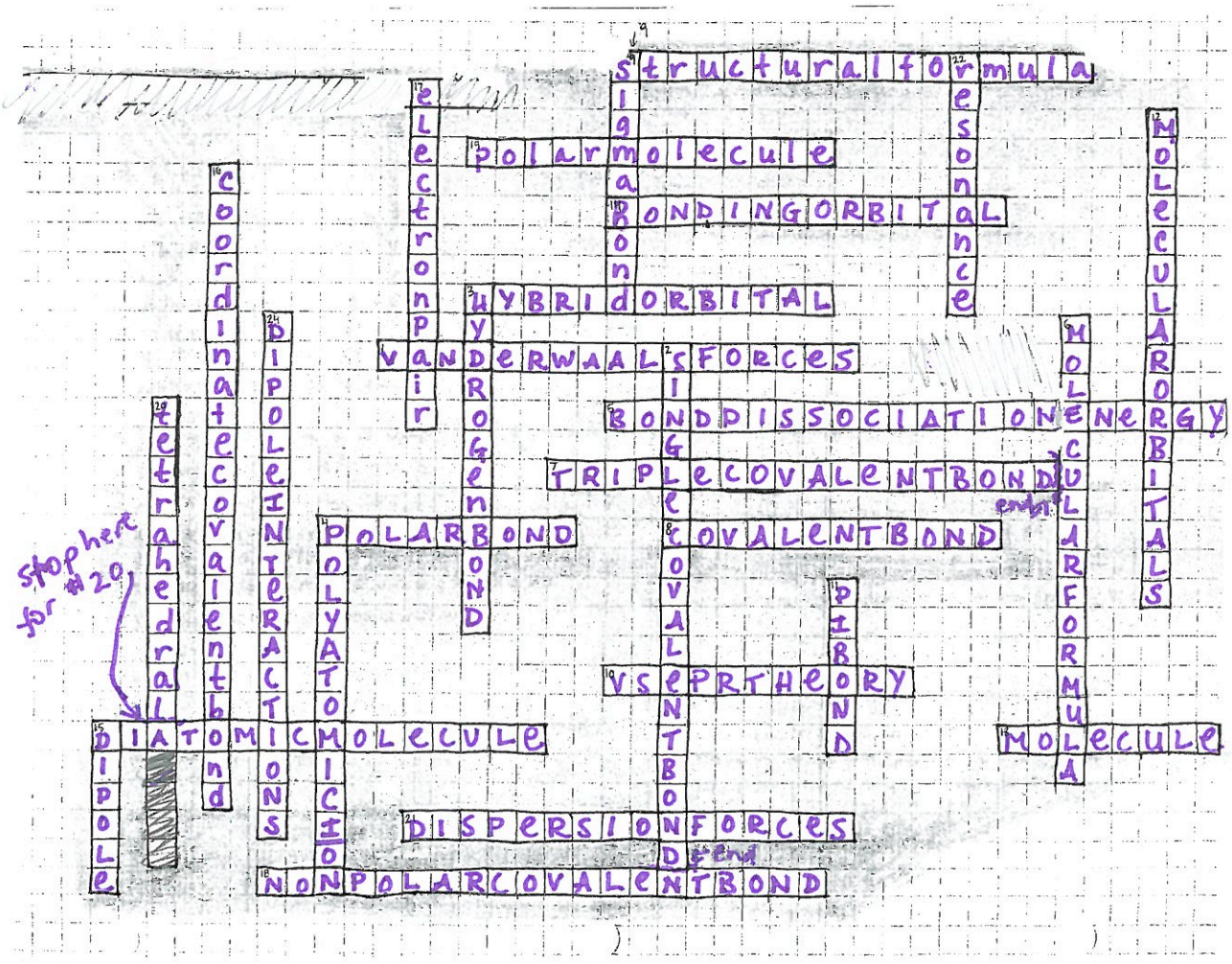
12.) Orbitals that apply to the entire molecule.

15.) Molecule that has 2 poles.

16.) Covalent bond in which 1 atom contributes both bonding electrons.

17.) Lone pair or a nonbonding pair. ← Lame

20.) Bond angle of 109.5 degrees.

22.) Type of structure that occurs when it is possible to draw 2 or more valid electron dot structures that have the same # of electron pairs for a molecule or ion.

24.) __ occur when polar molecules are attracted to one another.