

RATES OF REACTION

Section Review

Objectives

Describe how to express the rate of a chemical reaction
Identify four factors that influence the rate of a chemical reaction

Vocabulary

- rate
collision theory
activation energy
- activated complex
 - transition state
 - inhibitor

Part A Completion

In this completion exercise to check your understanding of the concepts and terms if one introduced in this section. Each blank can be completed with a term, short phrase, or number.

1. _____ measure the speed of any change that occurs within _____ time interval. Collision theory states that particles _____ when _____ collide, provided that they have enough _____.
2. _____
3. _____
4. _____ The rate at which a chemical reaction occurs is determined _____.
5. _____ energy barrier. The activation energy is the _____.
6. _____ The _____ by that reactants must have to be converted to _____.
7. _____ The _____ activation energy barrier, the _____ the reaction.
8. _____ catalysts help reactants overcome the activation barrier in a _____ number of ways. Two effective methods are to increase the _____.
9. _____ which the reaction is done or use a _____.
10. _____ Rates of reaction _____ also be increased by _____ the concentration of reactants.

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

1. An increase in temperature will increase the rate of a reaction.

12. A catalyst is considered as a reactant in a chemical reaction.
13. The speed of a reaction can be increased by increasing reactant concentration or decreasing particle size.
14. An enzyme is a biological catalyst.

Part C Matching

Match each description in Column B to the correct term in Column A.

Column A

Column B

- | | |
|-----------------------|---|
| 15. rate | a. synonym for an activated complex |
| 16. collision theory | b. speed of a change that occurs over time |
| 17. activation energy | c. substance that interferes with the action of a catalyst |
| 18. transition state | d. particles can react to form products when they collide, provided they have enough kinetic energy. |
| 19. activated complex | e. an unstable arrangement of atoms that forms momentarily at the peak of the activation energy barrier |
| 20. inhibitor | f. minimum energy that particles must have in order to react |

Part D Questions and Problems

Answer the following question and solve the following problem in the space provided.

21. An ice machine can produce 120 kg of ice in 24 hours. Express the rate of ice production in kg/h.
2. Which of the following will increase the rate of a reaction?
 - a. increase particle size
 - b. increase temperature
 - c. decrease concentration
 - d. add a catalyst