

Name

STURMAN KEY

★ Don't worry about those w/ an X in the box! Not on the test. Didn't cover it.

Chemistry - Ch 4/25 Review

You have a short amount of time to obtain a Bingo on the score card below! Mark of a square by finding a classmate who know the information within the square. Have your classmate initial the square. Notify your teacher when you have a Bingo, then continue to fill up your card. BE WARE - do not initial a square unless you know the answer and are prepare to present it to the class!

JJ Thomson hypothesized that the atom was a sphere of positive charge w/ electrons embedded.



<p>Knows the atomic number and mass of Cadmium</p> <p>112 Cd 48</p>	<p>Knows the difference between plum pudding and Rutherford's model of the atom</p> <p>Atom mostly space w/ dense nucleus</p>	<p>Knows at which elements on the periodic table are radioactive</p> <p>≥ 83</p>	<p>Knows the Beta decay of Terbium</p> <p>$^{159}_{65}\text{Tb} \rightarrow \beta^- + ^{159}_{66}\text{Dy}$</p>	<p>Knows who discovered and named the electron.</p> <p>JJ Thomson</p>
<p>Can define and give an example of an isotope</p> <p>$n^{\circ} \neq p^+$ $^{131}_{53}\text{I}$</p>	<p>Knows the direction of flow in the cathode ray tube</p> <p>- to +</p>	<p>Knows what Becquerel discovered</p> <p>discovered naturally occurring radioactivity</p>	<p>Knows the # protons, neutrons & electrons in Li⁺</p> <p>$^7_3\text{Li} + 3p^+ + 4n^{\circ} + 2e^-$</p>	<p>Can define atomic number and give an example</p> <p>#p⁺ Nitrogen N has 7p⁺ ∴ atom # is 7.</p>
<p>Knows the number of protons, neutrons and electrons in Cu²⁺</p> <p>$29p^+, 29n^{\circ}, 27e^-$</p>	<p>Can define amu and knows how many amu in protons, neutrons & electrons</p> <p>Compared to C-12</p>	<p>FREE SPACE</p>	<p>Knows what Chlorine-37 means</p> <p>isotope of Cl w/ 20n^o</p>	<p>Knows the charge and symbol for alpha, gamma and beta particles.</p> <p>-1β⁻ neg 4α⁺ positive 0γ neutral +1β⁺ positive</p>
<p>Knows who is credited with the Law of Constant Composition and what it says</p>	<p>Knows the charge of the cathode ray</p> <p>(-) neg</p>	<p>Knows the Alpha decay of Thorium</p> <p>$^{232}_{90}\text{Th} \rightarrow \alpha + ^{228}_{88}\text{Ra}$</p>	<p>Can list Dalton's 4 postulates (see below)</p>	<p>Can define and give an example of an ion</p> <p>Ca²⁺ ion + or - charged atom.</p>
<p>Knows the Beta decay of Rutherfordium</p> <p>$^{264}_{101}\text{Rf} \rightarrow \beta^- + ^{264}_{102}\text{Db}$</p>	<p>Can explain strong nuclear force</p> <p>Holds p⁺ together in the nucleus. Short range force.</p>	<p>Knows the # pro, neut and elect in Ba²⁺</p> <p>56p⁺ 56n^o 54e⁻</p>	<p>Can define atomic mass</p> <p>Weighted avg mass of naturally occurring isotopes in a sample.</p>	<p>Knows the Alpha decay of Uranium</p> <p>$^{238}_{92}\text{U} \rightarrow \alpha + ^{234}_{90}\text{Th}$</p>

Relative masses of elements are fixed in chemical compounds.