

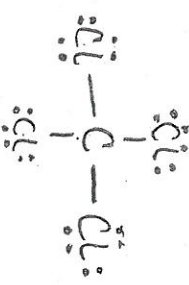
Name STURMAN KEY
 Date _____ Period _____

**Don't worry about coordinate bonds!*

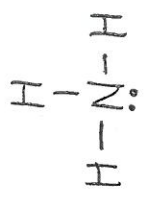
Activity: ELECTRON DOT STRUCTURES

Directions: Use the materials in your packet to create the electron dot structures for the following compounds. In the appropriate spaces below draw the structures with dashes indicating covalent bonds and dots as unshared electrons. Also, write the name of the compounds.

1. CCl_4 name carbon tetrachloride



2. NH_3 name nitrogen trihydride (Ammonia)



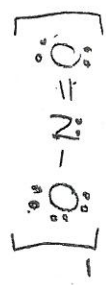
3. C_2H_2 name dicarbon dihydride (acetylene)
triple bond



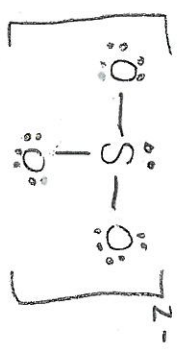
4. O_3 name trioxide (ozone)
double bond



5. NO_2^- name nitrite



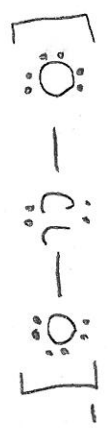
6. SO_3^{2-} name sulfite



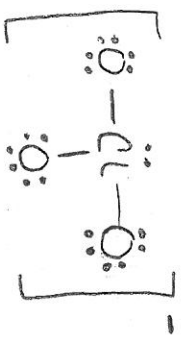
7. ClO^- name hypochlorite



8. ClO_2^- name chlorite
Hint: One coordinate bond.



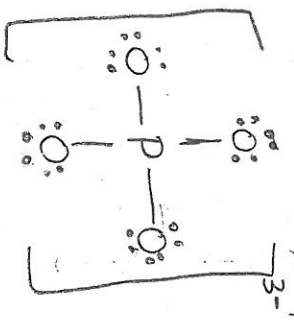
9. ClO_3^- name chlorate
Hint: Two coordinate bonds.



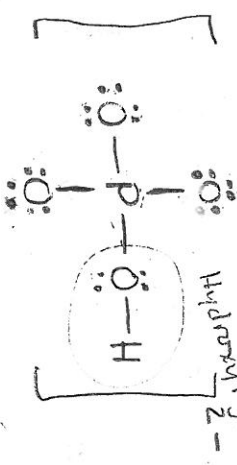
10. PO_3^{3-} name phosphite
Hint: Two coordinate bonds.



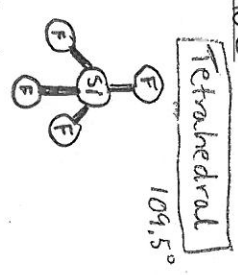
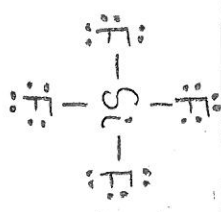
11. PO_4^{3-} name phosphate
Hint: Three coordinate bonds.



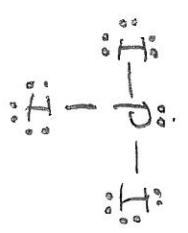
12. HPO_4^{2-} name hydrogen phosphate
Hint: One coordinate bond and hydrogen on outside.



13. SiF_4 name silicon tetrafluoride



14. PI_3 name phosphorus tri-iodide



15. SO_2 name sulfur dioxide
Hint: One coordinate bond.

