

## POLAR BONDS: EXTRA PRACTICE PROBLEMS

Compound \_\_\_\_\_  $|\Delta\chi|$  Bond Type, Structure & Bond Notation

1.  $\text{SeCl}_2$

\_\_\_\_\_

2.  $\text{SO}_2$

\_\_\_\_\_

3.  $\text{BCl}_3$

\_\_\_\_\_

4.  $\text{SiF}_4$

\_\_\_\_\_

5.  $\text{PI}_3$

\_\_\_\_\_

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Compound	$ \Delta\chi $	Structure & Bond Notation
1. $\text{SeCl}_2$	$=  2.55 - 3.16 $ $= .61 \text{ polar covalent}$	
2. $\text{SO}_2$	$=  2.58 - 3.44 $ $= .86 \text{ polar covalent}$	
3. $\text{BCl}_3$	$=  2.04 - 3.16 $ $= 1.12$ <p>polar covalent</p>	
4. $\text{SiF}_4$	$=  1.90 - 3.98 $ $= 2.08$ <p>Ionic</p>	<p>← No special notation if ionic or non-polar covalent</p>
5. $\text{PI}_3$	$=  2.19 - 2.66 $ $= .47$ <p>polar covalent</p> <p>* slightly since near .40</p>	